

Catalytic Self-Threading: A New Route for the Synthesis of Polyrotaxanes

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ABSTRACT: Main chain and branched polyrotaxanes have been synthesized in which polymerization and rotaxane formation occur simultaneously, due to the presence of the catalytically active self-threading macrocycle cucurbit[6]uril. Using monomers that contain stopper groups to prevent the catalytic macrocycle from noncatalytic threading, it was possible to prepare polyrotaxanes in high yields with molecular weights up to 39000. These polyrotaxanes are structurally perfect in the sense that exactly two macrocycles are threaded onto each structural repeat unit. Investigations into the polymerization mechanism have demonstrated that the catalyst cucurbit[6]uril is highly sensitive toward the structure of the monomers employed and a poorly designed monomer may result in complete inactivity. Features of the mechanism are discussed in some detail.

Introduction

One of the more recent polymer topologies that has received considerable attention is polyrotaxanes.^{1–6} Polyrotaxanes are supramolecular entities in which a macrocyclic compound is threaded onto a segment of a polymer main chain or side chain (Figure 1). The structure of the polymer threaded with the macrocycle and the strength of the noncovalent interaction between both determines if and to what extent the macrocycle can slide along the polymer main chain or side chain.^{1–5}

This latter design feature has been exploited to prepare thermoresponsive,^{7–9} photoresponsive,^{10–13} and pH-responsive^{14,15} polyrotaxanes with an ultimate application as triggerable macromolecular switching devices. The appropriate choice of macrocycle and its level of threading influences the solution-phase^{8,16–19} and solid-state^{16–22} behavior of the parent polymer, altering the solubility characteristics,^{8,16–19,23–25} phase behavior,^{8,16–19,23} and melt characteristics²⁶ as well as solution viscosities^{8,16–19,23,27} at significant levels. Achieving high levels of threading to protect and insulate conducting polymer backbones is an important goal in polymer electronics.²⁸ Biodegradable polyrotaxanes have been studied extensively for biomedical applications.^{14,29–34} As the sophistication of polyrotaxane syntheses increases the potential for new and unique applications increases concomitantly.

Methodologies applied to the synthesis of polyrotaxanes have improved quite dramatically since the initial statistical syntheses.^{1,35,36} The design of axel/macrocycle pairs with optimized steric and electronic fit,^{37–40} in conjunction with substantially more efficient and recognition-event-tolerant incorporation of stopper groups,⁴¹ has chiefly contributed to this synthetic progress. Most relevant to our investigations are main chain rotaxane architectures, which have five generic synthetic routes (Figure 2).

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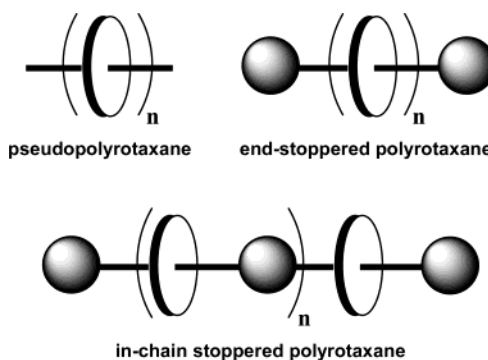


Figure 1. Definition of main chain polyrotaxanes.

Formation of a pseudopolyrotaxane via threading is strikingly simple in that a monomer is polymerized in the presence of a macrocycle, though in most cases there is little control over the number of macrocycles that can be incorporated.^{1,18,42–44} As an alternative to in situ polymerization, a preformed polymer can be threaded typically by mixing it with the macrocycle in solution under appropriate conditions.^{1,23,45–52} In some cases the pseudopolyrotaxane precipitates during this process. For all subsequent solution manipulations dethreading of the macrocycles becomes an issue.^{13,53,54} To prevent this from happening, blocking (“stopper”) groups have been incorporated either at both chain ends or as an integral part of the repeat unit along the polymer main chain classified as polyrotaxanes (Figure 2).⁵⁵ An alternative to stoppering after pseudopolyrotaxane formation^{31,56} is the use of prestoppered rotaxane monomers which are then polymerized^{57–62} or the copolymerization of self-assembled pseudorotaxane monomers with mono- or difunctional stopper groups.^{1,63–66} Clipping is conceptually the formation of macrocycles from linear segments in the presence of a polymer chain. This approach has been employed in various guises for the synthesis of rotaxanes,^{67,68} though it so far constitutes a rarity in the synthesis of polyrotaxanes.⁶⁹ Intrinsic to slipping is the presence of stopper groups along the polymer chain. Slipping was also first demonstrated in the synthesis of rotaxanes.^{36,70} The spatial demand of the blocking group in relation to the size of the cavity of

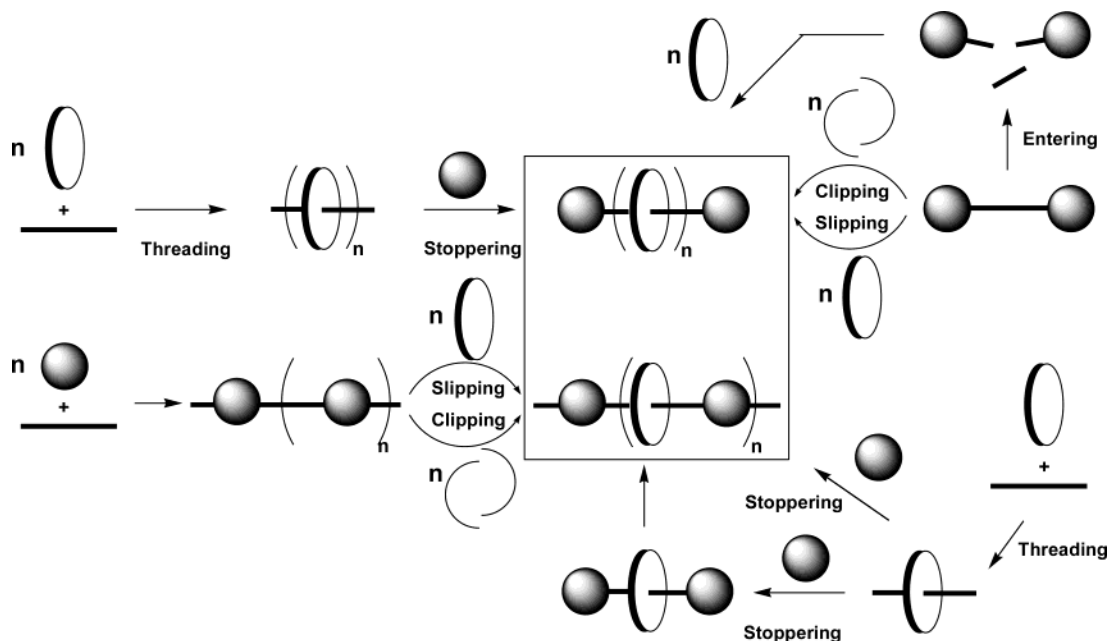


Figure 2. Generic routes for the synthesis of main chain polyrotaxanes.

the macrocycle to be threaded onto the polymer chain is crucial for successful slippage.^{71–74} Recently, polyrotaxanes have been obtained whereby the process of slipping combines polymerization with simultaneous threading using difunctional monomers.⁷⁵ The fifth strategy for the synthesis of main chain polyrotaxanes is entering. This approach relies on a polymer backbone that can be reversibly disassembled, with macrocycles becoming incorporated during the reassembly process (Figure 2).⁷⁶ In an elegant synthesis Takata et al. have demonstrated the validity of this concept by employing the pH dependence of disulfide/thiol scrambling equilibria in the formation of poly[3]rotaxanes.⁷⁷

None of the five strategies provide complete control over the number of macrocycles threaded. An exception is the polymerization of preformed rotaxane monomers, which is synthetically demanding¹ and positions each macrocycle within a narrowly defined space, thereby reducing the benefits of unusual behavior that arise from the interdependent polymer chain/macrocycle dynamics and their individual chemical characteristics.

Our motivation was to investigate a novel synthetic approach that may allow us to achieve ultimate control of threading without having to resort to preformed rotaxane monomers.^{1,16–19,21,60} Key to our methodology is a chemically and thermally exceptionally robust macrocycle that has been in scientific hibernation since its first synthesis by Behrend et al. in 1905.⁷⁸ The seminal work by Mock et al.^{79–89} in the 1980s brought this unusual molecule, to which the name of cucurbituril (now cucurbit[6]uril) was given in analogy to its shape being reminiscent of that of a gourd or pumpkin (Figure 3), to the attention of the supramolecular chemistry community.^{90–107}

Cucurbit[6]uril possesses a unique set of molecular recognition modes as a consequence of its rigid, nonadecacyclic shape, hydrophobic interior, and two hydrophilic oculi formed by a set of six equidistant carbonyl groups, each set aiming at a common focal point located approximately 10 Å away from the macrocycle's center of gravity.^{89,93} The opening of each portal is close to 4 Å in diameter through which small molecules can enter the interior.^{80,83,108} Modifications along the equatorial

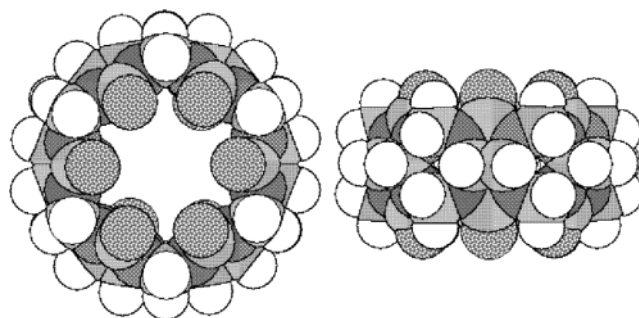


Figure 3. Space-filling representations of cucurbituril 1.

perimeter of cucurbit[6]uril have rendered it soluble in organic solvents.^{109,110} Larger members of the cucurbituril family have been explored as new hosts for supramolecular assemblies.^{105,111–114} Cucurbit[6]uril itself binds strongly to a variety of main-group^{115–118} and transition-metal^{119,120} ions and to aliphatic^{80,83–85,121,122} and aromatic mono- and diammonium ions,^{80,83–85,123–125} primarily through ion–dipole interactions. Although the interior is hydrophobic toward the center of the cavity, it gradually becomes more hydrophilic toward the carbonyl groups.¹⁰⁷ The latter are part of the cyclic urea moieties which are responsible for the highly hygroscopic nature of cucurbit[6]uril.⁹⁷ The property that distinguishes cucurbit[6]uril from any of the typical macrocycles (cyclodextrins, crown ethers, cyclophanes) is its remarkable ability to catalyze 1,3-dipolar cycloadditions in a regioselective manner (Figure 4).^{81,86,103,106}

Our idea was to extend the catalysis shown by cucurbit[6]uril to the synthesis of polyrotaxanes by a conceptually simple extension in which azide and alkyne monomers of the A₂ and B₂ types replace those chosen in the original detailed catalytic studies conducted by Mock et al.^{81,86,89}

Results and Discussion

1. First-Generation Monomers. The design of the monomers was based on the features first identified by Mock et al.^{81,86,89} The best spacing between the alkyne substituent and the ammonium ion was found to be a

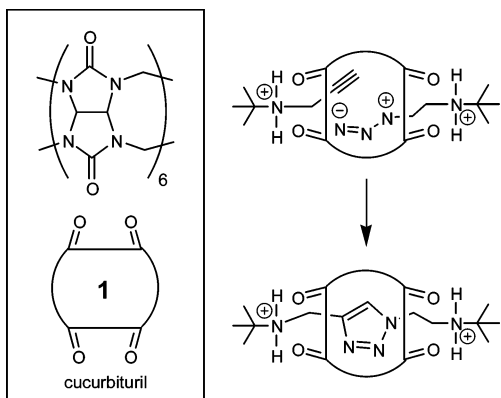


Figure 4. 1,3-Dipolar cycloaddition catalyzed by cucurbituril.

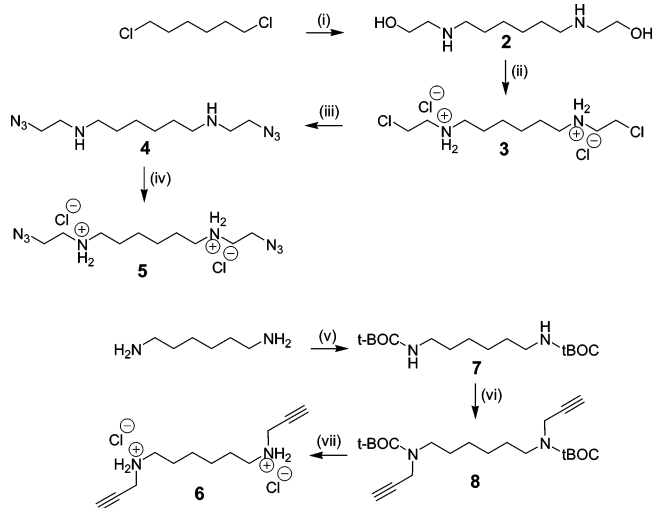


Figure 5. Synthesis of aliphatic A₂-type and B₂-type monomers **5** and **6**. Conditions and reagents: (i) 2-aminoethanol (neat), 20 min at 120–130 °C and then 6 h at 150–160 °C, 57%; (ii) SOCl₂, 60 °C, 100 min, 65%; (iii) NaN₃, H₂O, 75 °C, 16 h, 60%; (iv) 1 N HCl in Et₂O, room temperature, 84%; (v) [(CH₃)₃CO]₂O, dioxane, room temperature, 48 h; (vi) BrCH₂CCH, DMF, NaH, room temperature, 18 h; (vii) 1 N HCl in diethyl ether, room temperature, 4 h.

methylene group,⁸⁹ whereas for the azido substituent an ethylene spacer provides the best distance to the ammonium group.⁸⁹ As the ideal spacer length to link two ammonium groups, we chose a simple hexamethylene chain because the high affinity of cucurbit[6]uril to hexamethylenediammonium ions ensures essentially 100% encapsulation of the monomer by cucurbit[6]uril.⁸⁴ The final polymer structure would be a pseudopolyrotaxane in nature, which we hoped could be triggered to undergo dethreading upon a change in pH.

We prepared monomers **5** and **6** as outlined in Figure 5. Monomer **5** was synthesized in four steps by first N-alkylation of ethanolamine with 1,6-dichlorohexane to obtain compound **2** followed by chlorination of the hydroxy groups with SOCl₂. Conversion of the resulting dichloride **3** to its hydrochloride salt form using NaN₃ furnished the neutral diazido species **4**, which is directly converted into its hydrochloride salt **5** obtained in 17% overall yield. Monomer **6** was synthesized starting from 1,6-hexanediamine which was converted into its *tert*-butoxycarbonyl (*t*-BOC) protected form **7**¹²⁶ followed by alkylation with propargyl bromide.¹²⁷ Deprotection of **8** and its conversion into the corresponding hydrochloride salt produced **6** in 36% yield. Cucurbit[6]uril was

Table 1. Reaction Conditions for the Attempted Polymerization of Monomers **5** and **6** in the Presence of **1**

entry	[5]	[6]	[1]	<i>t</i> /h	<i>T</i> /°C	[5 + 6]/[TrH]	DP _n ^a
A	1.00	1.00	1.00	24	20	16/0.0	0.0
B	1.00	1.00	2.00	192	20	16/0.0	0.0
C	1.00	1.00	3.00	120	20	16/0.0	0.0
D	1.00	1.00	2.00	384	90	16/1.0	1.0
E	1.00	1.00	3.00	384	90	16/0.8	0.9
F	1.00	1.00	4.00	384	90	16/0.7	0.8
G	1.00	1.00	0	384	90	16/0.8	0.9

^a DP_n = degree of polymerization defined by the ratio of the integral for the sum of both encapsulated (0.75 and 0.45 ppm) and free β- and γ-methylene (1.65 and 1.35 ppm) protons of each monomer (**5** + **6**) divided by the integral for the triazole proton (H_{Tr}, 8.55 ppm).

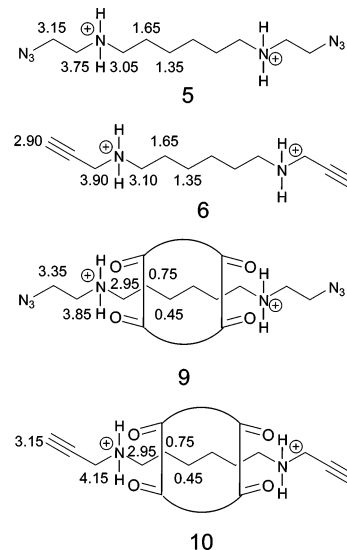


Figure 6. ¹H NMR chemical shift assignment for monomers **5** and **6** and [2]pseudorotaxanes **9** and **10**.

synthesized according to the procedure by Behrend et al. with minor modifications.⁷⁸

2. Polymerization of First-Generation Monomers. Polymerization of monomers **5** and **6** in the presence of cucurbit[6]uril was then attempted by dissolving the cavitand in 6 N HCl (2 equiv) before **5** (1 equiv) and **6** (1 equiv) were added. After 8 days at room temperature, using the conditions we had established prior to the polymerization for the rotaxane formation (Table 1, entry B), the ¹H NMR spectrum of the crude product mixture only showed signals corresponding to [2]pseudorotaxanes **9** and **10** derived from monomers **5** and **6** encapsulated by cucurbit[6]uril (Figure 6). The ¹H NMR spectrum was completely devoid of any aromatic resonances, indicating the absence of any triazole protons. This was unexpected, though a possible explanation for this outcome is the slow dissociation rate constant (10⁻⁴ s⁻¹)⁸³ in conjunction with the large complexation constant *K*_{ass} between cucurbit[6]uril and monomers **5** and **6**, respectively, of approximately 10⁶–10⁷ M⁻¹.⁸³ These numbers translate into only minute quantities of free cucurbit[6]uril being available for catalysis. Considering that the overall rate of the cycloaddition reaction is already slow to begin with,⁸⁶ it becomes clear why triazole formation has not been observed. The reaction was repeated by adding a further equivalent of cucurbit[6]uril, ensuring that excess cavitand was available (Table 1, entry C), but again no catalysis was observed. Further variations in the reaction conditions such as different stoichiom-

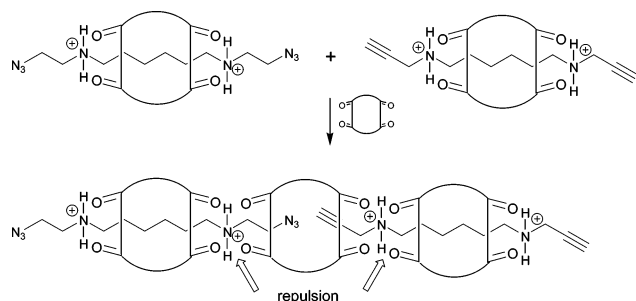


Figure 7. Suggested explanation for the suppressed catalytic activity of cucurbituril **1** in the presence of monomers **5** and **6**.

etries of reactants, longer reaction times, and higher reaction temperatures were investigated but to no avail (Table 1).

Triazole formation was observed only for reactions at elevated temperatures and prolonged reaction times. The highest degree of polymerization (DP_n) was achieved under the conditions shown in Table 1 (entry D). By comparing the integral for the central methylene protons of the monomers (free and encapsulated) with that of the triazole proton, one can calculate the DP_n of this reaction, which is merely 1. This value suggests that under the conditions explored it is not possible for more than two monomers to react together, before the reaction somehow shuts down. Thermal instability of the azide groups in monomer **A** is likely to be one contributing factor. Furthermore, the presence of excess cucurbit[6]uril (Table 1, entries E and G) could reduce the statistical chances for azido groups to encounter alkyne moieties due to encapsulation of those cucurbit[6]urils not threaded onto **5** or **6**. This would explain that the same number of triazole rings was formed regardless of the monomers being heated on their own or in the presence of excess cucurbit[6]uril (Table 1, entries E–G). A small but significant increase in the number of triazole rings formed when only 2 equiv of cucurbit[6]uril was present (Table 1, entry D). We rationalize this result by assuming a relative increase in the accessibility of functional groups in general together with some catalytic contribution from cucurbit[6]uril. For steric reasons though, the conformation of the ternary complex essential for catalytic activity cannot be adopted at any further stage of the reaction. This hypothesis is illustrated in Figure 7.

3. Pseudopolyrotaxane Formation via Post-threading. To test our hypothesis regarding the structural factors that were responsible for the catalytic inactivity of **1**, we prepared model polymer **11** which contains protonated amine loci spaced very similar to those found in monomers **5** and **6** as well as the resulting triazole product. Nylon 6/6 was reduced according to a procedure by Schulz et al.¹²⁸ to obtain poly(iminohexamethylene) **11** ($M_n = 20000$, PDI = 2.7), in 66% yield (Figure 8).

NMR-scale experiments were performed by dissolving **1** in DCl/D₂O (20% w/w) first, before **11** was added. The initial threading experiment was carried out at room temperature using ratios of **1** to polymer chain repeat unit from 0.1:1 to 5:1. After 9 days at room temperature levels of threading for all stoichiometries were still below 10%. Heating identical samples to 90 °C both accelerated the threading process and substantially increased the levels of threading.⁴⁸ The results are summarized in Table 2.

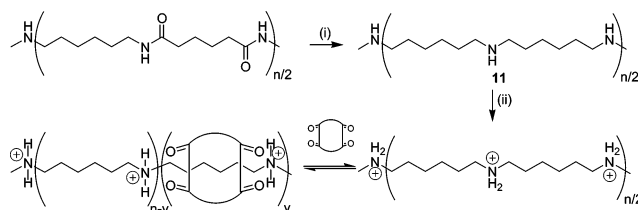


Figure 8. Synthesis of **11** followed by threading with cucurbituril **1**. Conditions and reagents: (i) $BH_3 \cdot DMS$, THF, reflux, 66%; (ii) DCl/D₂O (20% w/w), 20 and 90 °C.

Table 2. Reaction Conditions and the Degree of Threading for Cucurbituril **1** Threaded onto **11** in DCl/D₂O (20% w/w) at 20 and 90 °C

<i>t</i> /h	[1]/[<i>n</i>] ^a					
	5.0/1	1.0/1	0.5/1	0.3/1	0.2/1	0.1/1
At 20 °C						
48	0.02	0.02	0.05	0.06	0.05	0.04
72	0.05	0.05	0.06	0.07	0.06	0.06
216	0.08	0.07	0.07	0.08	0.08	0.06
At 90 °C						
48	0.33	0.23	0.26	0.22	0.18	0.10
144	0.42	0.41	0.41	0.30	0.19	0.10
216	0.47	0.43	0.44	0.30	0.20	0.11
384	0.49	0.45	0.46	0.30	0.20	0.11
456	0.50	0.45	0.46	0.30	0.20	0.10

^a[1]/[*n*] = degree of threading; *n* = number of repeat units.

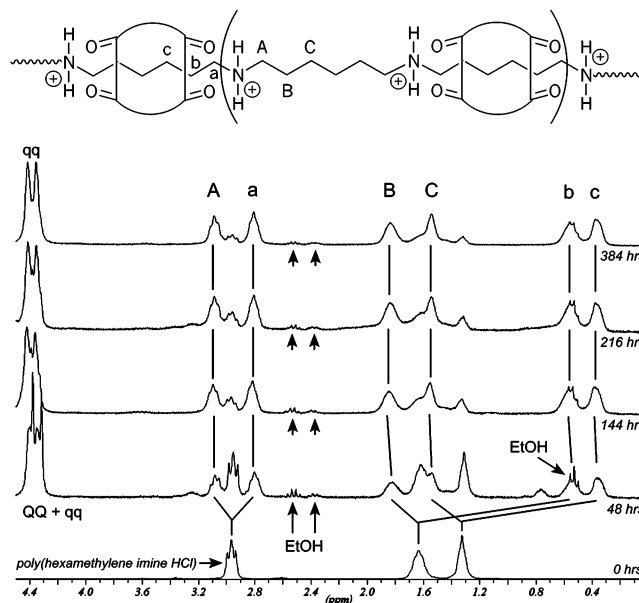


Figure 9. Time dependence of postthreading followed by ¹H NMR at 90 °C in DCl/D₂O (20% w/w) with a molar ratio of cucurbituril to hexamethyleniminium repeat unit of 0.5:1.0. Arrows indicate traces of an ethanol inclusion complex with cucurbituril (ethanol was used as cosolvent during recrystallization) decreasing with time progressively replaced by polymer chain repeat units.

The progress of the threading reaction was monitored at given time intervals by ¹H NMR (Figure 9). The degree of threading is defined as the ratio of **1** per hexamethyleniminium repeat unit. Figure 9 corresponds to a threading experiment carried out at 90 °C in DCl/D₂O (20% w/w) with a molar ratio of cucurbit[6]uril to repeat unit of 0.5:1.0 with ¹H NMR spectra recorded after 48, 144, 216, and 384 h. For comparison the spectrum of the unthreaded polymer backbone has been included. Upon threading, the original three resonances of the α-, β-, and γ-methylene groups of poly(iminium hexameth-

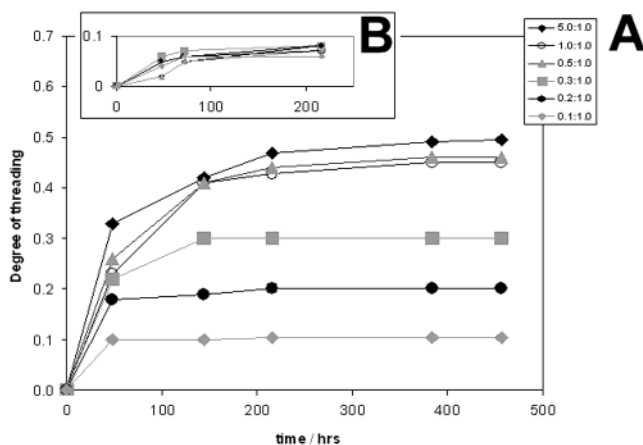


Figure 10. Time dependency of the threading of cucurbituril onto poly(iminium hexamethylene) chloride at 90 °C (graph A) and 20 °C (graph B, inset) in DCl/D₂O (20% w/w). The molar ratios of cucurbituril to repeat unit used are 0.1:1.0, 0.2:1.0, 0.5:1.0, 1.0:1.0, and 5.0:1.0.

ylene) chloride are split into an additional two sets of resonances (Figure 9, A–C and a–c). The integrals of the new resonances increase over time to the same extent as the resonances for unthreaded hexamethylene segments decrease. Resonances coded a–c have been assigned to threaded hexamethylene segments, on the basis of complexation studies of **1** with the model compound 1,6-diaminohexane dihydrochloride salt in DCl/D₂O (20% w/w) and the data obtained from [2]-pseudorotaxanes **9** and **10** (see Figure 6), with further support by earlier work in aqueous formic acid carried out by Mock et al.⁸³ Resonances A–C belong to unthreaded hexamethylene segments next to threaded repeat units along the polymer backbone. A downfield shift is experienced by all protons likely due to the proximity of the magnetically anisotropic carbonyl groups of **1**. The assignments were confirmed by repeating the same threading experiment with a 5-fold excess of cucurbit[6]uril since under these conditions essentially all precursor resonances disappeared and only two new sets of resonances (A–C and a–c) remained. ¹H NMR data from the threading of 1 equiv of cucurbit[6]uril onto bis(6-aminoethyl)amine trihydrochloride are also consistent with this interpretation.

The degree of threading was calculated from the ratio of threaded to nonthreaded methylene protons and was cross-checked by comparing the ratio of threaded methylene protons to cucurbit[6]uril protons. In the example shown (Figure 9) 46% of all repeat units were threaded by cucurbit[6]uril after 384 h. Threading is also indicated by a small downfield shift in the resonances of cucurbit[6]uril (Figure 9, QQ, free **1**; qq, threaded **1**), also proportional to the degree of threading. Figure 10 illustrates the very slow kinetics of the threading process at 20 °C (inset B) and 90 °C. The data suggest that the system is only close to equilibrium after heating for almost 400 h at 90 °C for all ratios of cucurbit[6]uril to polymer repeat unit investigated. In the presence of 0.1 or 0.2 equiv of **1** per repeat unit the final degree of threading is very close to the theoretical maximum given the experimental margin of error of 1–2%. At a ratio of 0.5:1 the degree of threading reaches a value of 46%, close to but significantly below the theoretical limit. A supply of equimolar amounts of cucurbit[6]uril had little impact on the maximum degree of threading, though by using a 5-fold excess of **1** our

data suggest that a degree of threading of essentially 50% was reached, suggesting the existence of an upper limit. We believe that, through the limit found for the degree of threading (50%) and the clear splitting of the original ¹H NMR signals of the hexamethyleniminium repeat units into two new sets of peaks of equal intensity, this can best be rationalized on symmetry grounds by invoking the formation of a pseudopolyrotaxane with alternating encapsulated and uncomplexed repeat units. Further evidence to this effect was provided by encapsulation studies between cucurbit[6]uril and bis(6-aminoethyl)amine trihydrochloride, the triamine being used as a “dimer” model for the polymer backbone. A 2-fold excess of cucurbit[6]uril led to 45% encapsulated hexamethyleniminium segments, a value close to that found for the threaded polymer. Other substructures were also considered, in a vein similar to that of recent ¹H NMR investigations by Hodge et al.,⁵⁰ though we have concluded that an alternating sequence of threaded and nonthreaded repeat units is most consistent with our experimental data.

Intriguingly, at the maximum level of threading of 30% for a ratio of 0.3:1 (Table 2, Figure 10), the ¹H NMR spectra showed three resonances of almost equal intensity corresponding to the α -methylene protons (in appearance somewhere between the ¹H NMR spectra after 48 and 144 h in Figure 9) as expected for a degree of threading of close to 1/3. The chemical shift region (1.2–2.0 ppm) of the β - and γ -methylene protons showed six overlapping resonances of similar intensity, which indicates a high level of order along the polymer backbone. In fact there are only three possible threading arrangements that can explain this observation, though unfortunately the dynamic nature of the pseudopolyrotaxane and the general level of peak broadening have made it impossible so far to distinguish between them.

The slow threading kinetics overall can be explained through a combination of factors. By far the most important contribution arises from the strong association constant of cucurbit[6]uril to the protonated hexamethylene-spaced repeat unit in combination with a slow dissociation rate constant as has been shown for a model repeat unit.^{83,84} Therefore, efficient “hopping” of cucurbit[6]uril along the polymer backbone requires a high activation energy as illustrated by the differences in threading kinetics at 20 and 90 °C. Further support for this interpretation has been provided by Harada et al., who have related slow threading kinetics to strong ionic interactions between the macrocycle and polymer chain.⁶² Similar behavior has been observed also by Kim et al. in the threading of cucurbit[6]uril onto oligoviologenes.¹²⁹ Wenz et al. studied the postthreading of α -CDs onto ionene-6,10 and calculated the observed threading kinetics successfully by employing a hopping mechanism as the model.¹³⁰ Their pseudopolyrotaxane system also required elevated temperatures and long reaction times (120 h at 80 °C in water) to reach 50% as their highest degree of threading. Furthermore, we argue that threading kinetics in our system are further retarded by cucurbit[6]uril moieties having to queue especially toward higher threading levels. Queuing takes place in the formation of pseudopolyrotaxanes as has been demonstrated by Parsons et al.^{52,131} As expected the effect of queuing becomes more significant also in our system for higher levels of threading as reflected in the slowing of the threading progress with increasing time (Figure 10).

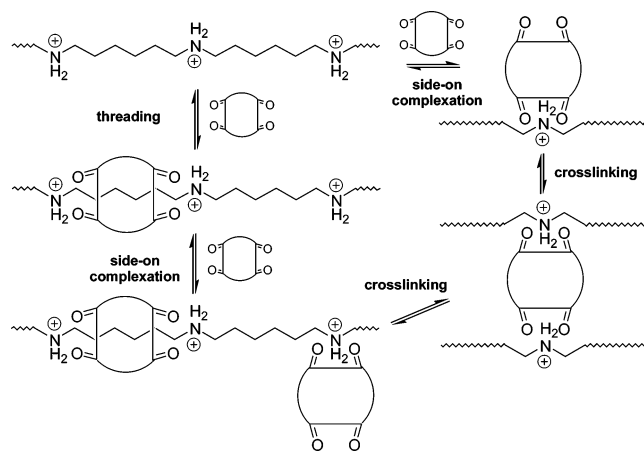


Figure 11. Dynamic equilibria that are part of the complex threading process of cucurbit[6]uril onto linear poly(iminium oligoalkylene)s.

Possessing two portals which can also bind to ammonium ions without the need for forming an inclusion complex, unthreaded cucurbit[6]uril may interfere in the hopping process through side-on complexation and formation of dynamic physical cross-links. Figure 11 is therefore a more accurate representation of the competing modes of complexation that have to be considered to describe the present system in sufficient detail. If one assumes that the kinetics of the side-on complexation are significantly faster than the K_{diss} of the chain hopping, then the effect will be small. Additional experiments will be required to establish the contributions of each of these modes of complexation.

4. Second-Generation Monomers. These threading studies have given us vital information on the consequences of spacer length to the complexation behavior of cucurbit[6]uril. Clearly, the hexamethylene spacer chosen in the design of monomers **5** and **6** is too short to simultaneously allow two individual cucurbit[6]uril portals to complex to the same ammonium ion. Furthermore, the dissociation rate constant is too small to produce sufficient numbers of cucurbit[6]uril available to form ternary complexes required for catalysis. At this juncture we could investigate the effect of the alkyne and azide substrates on the kinetics of the 1,3-dipolar cycloaddition and try either to enhance its rate or, and this appeared to be a more rewarding avenue, to reduce the association rate constant between cucurbit[6]uril and each monomer, in effect the design of new monomers with a much reduced affinity toward encapsulation by cucurbit[6]uril.

The most prudent approach would be to rule out the possibility of cucurbit[6]uril threading onto any new monomer through the introduction of stopper groups.¹³² Once again the seminal work by Mock et al. provided us with the necessary guidance. Groups bulkier than cyclohexyl rings¹⁰⁸ and disubstituted aromatic six-membered rings with the exception of 1,4-disubstituted ones are not included in the interior of cucurbit[6]uril.^{80,83,84} We identified 2,4-bis(chloromethyl)-1,3,5-trimethylbenzene (**12**), from which we could synthesize both the bisalkyne and the bisazido monomers (Figure 12).

N-[2,4,6-Trimethyl-3-[(2-propynylamino)methyl]-benzyl]-2-propyn-1-amine dihydrochloride (**14**) was synthesized via *N*-alkylation of propargylamine to yield crude **13** in 100% yield and directly converted to its HCl salt in 86% overall yield. The steric demand of the ortho-

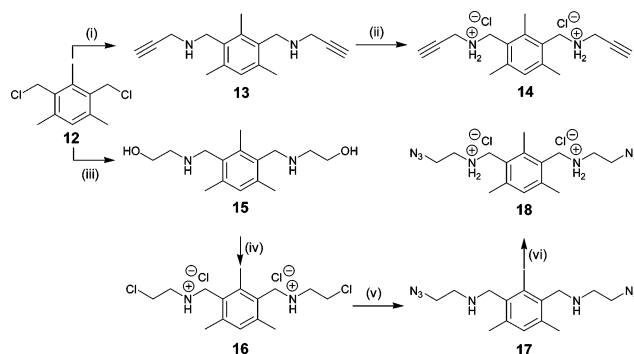


Figure 12. Synthesis of monomers **14** and **18**. Conditions and reagents: (i) $\text{NH}_2\text{CH}_2\text{C}\equiv\text{CH}$, neat, $0^\circ\text{C} \rightarrow \text{room temperature}$, 16 h, 100% (crude); (ii) 1 N HCl in Et_2O , room temperature, 86%; (iii) $\text{NH}_2\text{CH}_2\text{CH}_2\text{OH}$, neat, 5 h, $150\text{--}160^\circ\text{C}$, 66%; (iv) SOCl_2 , CHCl_3 , 5 h, room temperature, 70%; (v) NaN_3 , H_2O , 75°C , 16 h, 87%; (vi) 1 N HCl in Et_2O , room temperature, 96%.

positioned methyl groups on the ring was sufficient to suppress the possible second nucleophilic attack of the alkylated propargylamine toward **12**. This steric “protecting-group” approach proved equally useful in the preparation of *N*-(2-azidoethyl)-*N*-(3-[(2-azidoethyl)-amino]methyl)-2,4,6-trimethylbenzylamine dihydrochloride (**18**). Starting material **12** was alkylated first with ethanolamine, subsequently treated with thionyl chloride to yield **16**, and isolated as its dihydrochloride salt. Use of excess sodium azide led to the formation of **17** sufficiently pure to prepare dihydrochloride salt **18** directly (39% overall yield).

5. Polymerization of Second-Generation Monomers. With **14** and **18** in hand we repeated our initial attempts to trigger cucurbit[6]uril into catalytic activity (see Figure 7).¹³² In a typical polymerization **1** was dissolved in 6 N HCl followed by addition of first **14** and then **18**. For workup, each solution was precipitated into a mixture of acetone and ethanol to yield a white solid. Excess **1** was removed by dissolution of the polymer in hot water (80°C) followed by filtration. The final polymer was obtained as a colorless film. Polymerization conditions are summarized in Table 3.

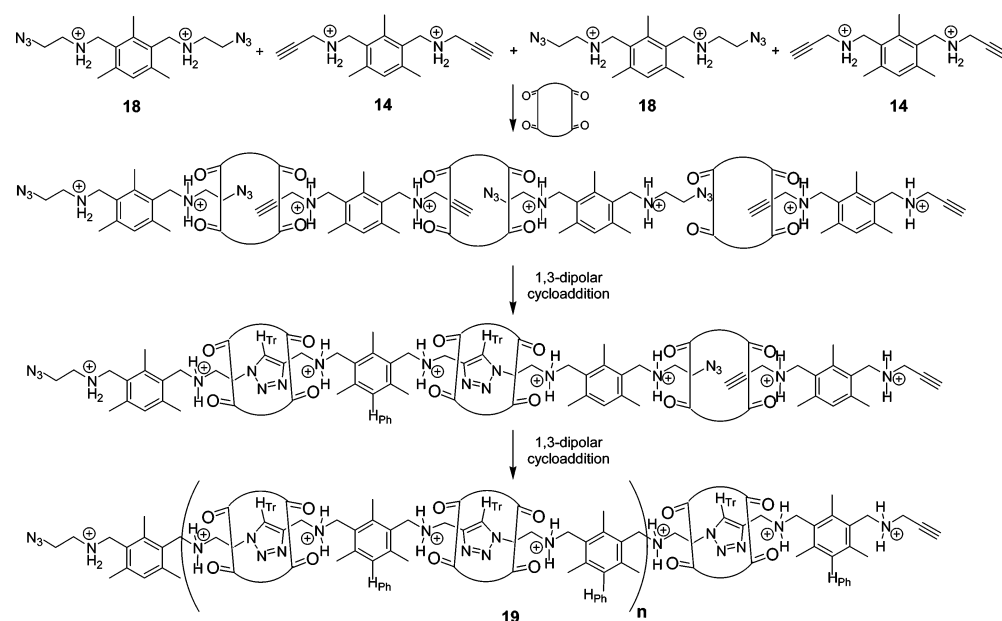
^1H NMR became the most valuable and revealing tool for the structural characterization of the polymerization products. The structure of the polymer that we expected to form (**19**) is shown in Figure 13. The most informative proton resonances are those of the triazole proton (H_{T}) and the phenyl proton (H_{Ph}). Both resonances appear in a region of the ^1H NMR spectrum not occupied by any other species as shown in Figure 14. Our ^1H NMR assignments of **19** were ascertained through a comparison with ^1H NMR data of model rotaxanes. To this end [2]rotaxane **20** and [3]rotaxane **21**¹⁰⁶ were prepared by reacting stoichiometric amounts of azide **22** with alkyne **23** or **14**, respectively (Figure 15).

Figure 14 illustrates that the triazole proton is positioned at around 6.5 ppm whereas the phenyl proton is detectable close to 7.2 ppm. The assignment of the model compounds is reflected in the resonances found for polyrotaxane **19** (^1H NMR of entry A in Table 3). The presence of the triazole proton in itself demonstrates that cucurbit[6]uril has acted as a catalyst joining together monomers **14** and **18** through a newly formed 1,3-disubstituted triazole. By taking another look at structure **19** in Figure 13, one realizes that each repeat unit of the polymer contains two triazole protons and two phenyl protons. Integration of the correspond-

Table 3. Reaction Conditions and Molecular Weight Data for the Polymerization of Monomers **14 and **18** in the Presence of **1** Dissolved in 6 N HCl^a**

entry	[14]/molar equiv	[18]/molar equiv	[1]/molar equiv	t/h	T/°C	yield/%	M _n (¹ H NMR)	DP _n
A	1.00	1.00	2.00	48	20	68	13000	5
B	1.00	1.00	2.00	144	20	82	25000	10
C	1.00	1.00	2.00	336	20	75	30000	11
D ^b	1.00	1.00	2.00	192	20	79	21000	8
E	1.00	1.00	2.00	72 + 24	20 + 40	75	28000	11
F	1.00	1.00	2.00	120	60	68	39000	15
G	1.00	1.00	2.00	144	80	91	29000	11
H	1.00	1.00	2.00	144	90	86	18000	7
I	1.00	1.00	2.00	336	90	79	25000	10
J	1.00	1.00	2.25	144	20	73	7000	3
K	1.12	1.00	2.36	144	20	72	12000	5
L	1.20	1.00	2.48	144	20	71	12000	5
M	1.00	1.00	2.25	144	90	84	12000	5
N	1.00	1.00	1.00	192	90	ni	3400	0.5
O	1.00	1.00	0	192	20	100	na	na
P	1.00	1.00	0	192	90	ni	300	0.5

^a DP_n = degree of polymerization; na = not applicable; ni = not isolated. See the text and Figure 3 for the calculation of molar mass and DP_n. ^b Solvent: 0.2 M Na₂SO₄.

**Figure 13.** Synthesis of a “perfect” main chain polyrotaxane (**19**) via catalytic self-threading. See Table 3 for reaction conditions.

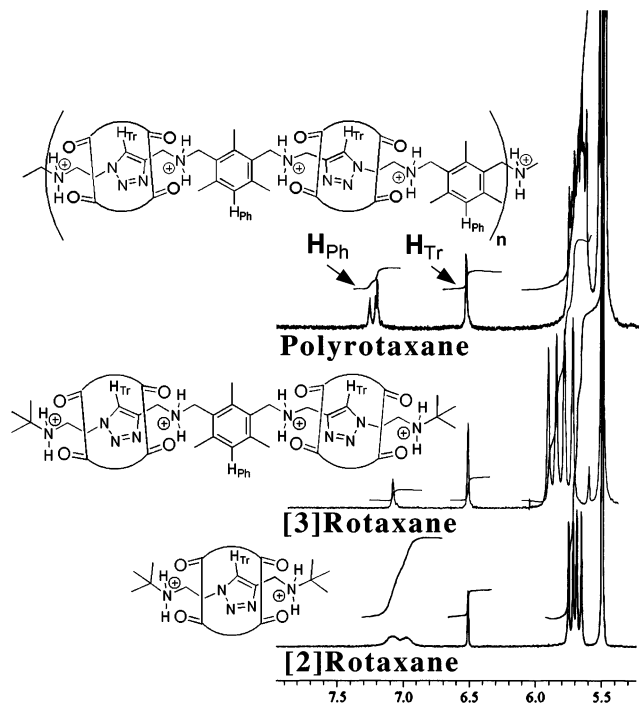
ing resonances in the ¹H NMR spectra allows us to determine the degree of polymerization according to Carothers' equation. An infinite polymer chain would have been formed when the ratio of the peak areas of integration of H_{Tr} and H_{Ph} was equal to 1. In the example shown in Figure 14 the ratio was calculated to be 0.9, which is equivalent to a degree of polymerization of 5. With the molecular weight of a repeat unit (see Figure 13) close to 2600, the number-average molecular weight of the polymer therefore is about 13000 and the number-average degree of polymerization (DP_n) is 10. Further evidence that the presence of cucurbit[6]uril is pivotal for polymerization to take place is found by reducing the number of equivalents of cucurbit[6]uril from 2 to 1 and by carrying out the reaction without cucurbit[6]uril all together. Equimolar amounts of monomers **14** and **18** and catalyst **1** (Table 3, entry N) brought about the expected formation of the [2]rotaxane. In the absence of **1**, monomers **14** and **18** show no sign of undergoing 1,3-dipolar cycloaddition at room temperature preceded by complete recovery of starting materials (Table 3, entry O). However, cycloaddition chemistry takes place when both monomers are heated for 9 days at 90 °C, in which case an

average of one triazole ring is formed for every set of **14** and **18** present.

We found that increasing the polymerization time at room temperature led to an increase in the molecular weight of the polyrotaxane of up to 30000 after 336 h in good yields (Table 3, entries A–C). Entries E–I in Table 3 show the effect of an increase in temperature on the number-average molecular weight. The data suggest an optimum is found at 60 °C for 5 days (Table 3, entry F). Longer heating periods and higher temperatures led to substantially lower molecular weights. We suspect the polymerization to be very sensitive toward changes in reaction conditions, in particular temperature. A possible explanation for this behavior could be the very crowded nature of the polymer backbone that is being formed, hypothesizing that the catalytic self-threading process is reversible. Consequentially, a higher temperature brings the system closer to its ceiling temperature. The reason for lower molecular weights at 20 °C instead of 60 °C could be explained by the slow rate of polymerization that has not allowed us to reach equilibrium conditions within the polymerization time intervals studied.

Table 4. Reaction Conditions and Molecular Weight Data for the Attempted Polymerization of Monomers **6 and **14** in the Presence of **1** Dissolved in 6 N HCl^a**

entry	[6]/molar equiv	[14]/molar equiv	[1]/molar equiv	t/h	T/°C	yield/%	M _n (¹ H NMR)	DP _n
A	1.00	1.00	1.00	168	20	100	monomers	0
B	1.00	1.00	2.00	168	20	75	dimer	0.5
C	1.00	1.00	3.00	168	20	ni	monomers	0
D	1.00	1.00	3.00	552	20	93	30% dimer	0.2
E	1.00	1.00	3.00	192	60	80	dimer	0.5

^a DP_n = degree of polymerization; ni = not isolated.**Figure 14.** ¹H NMR overlay of the triazole and phenyl proton regions of polyrotaxane **19** (Table 3, entry A) and model compounds [2]rotaxane **20** and [3]rotaxane **21**.

Some additional evidence can be extracted from another set of experiments (Table 3, entries J–N) designed to investigate the influence of polymerization stoichiometry. A reduction in the number-average molecular weight by 50% was found when one of the monomers was used in subequimolar amounts (Table 3, entries K and L), thereby upsetting the required stoichiometric proportions. This is the trend one would expect from a step-growth polymerization process. More interestingly though, we also found that by using 2.25 equiv instead of 2.00 equiv of **1** that *M_n* of the polyrotaxane was reduced by a factor of almost 4 from 25000 to 7000 at room temperature (Table 3, entries B and J) and to 2/3 of the value from 18000 to 12000 at 90 °C (Table 3, entries H and M). If cucurbit[6]uril is only taking part in this polymerization as a catalyst, then excess amounts of it should not affect the degree of polymerization. Excess **1** may slow the process, though even at 90 °C an excess of cucurbit[6]uril results in reduction of the molecular weight from 18000 to 12000 (Table 3, entries H and M). The question arises of whether an excess of cucurbit[6]uril could be responsible for shifting the polymerization equilibrium. At present, our data suggest that such a possibility exists, but we have insufficient data to draw any firm conclusion.

Further characterization of the polyrotaxanes was carried out employing MALDI-TOF mass spectrometry. Obtaining satisfactory spectra was only met with limited success as depicted in Figure 16. The appearance

of all spectra was characterized by a wide range of molar masses shaped into broad “humps” with little analytical value. MALDI-TOF spectra of polyrotaxane **19** at the high end of the molecular weight scale (Table 3, entries C and E–G) did not differ substantially from the spectrum shown. A little comfort can be derived from the spacing of those clusters though, as they correspond to a molecular weight of ~1350, the approximate value for a cucurbit[6]uril-threaded segment of the polymer backbone. We also attempted to carry out gel permeation chromatography on these polyrotaxanes. To render **19** soluble in a polar organic solvent such as DMF, we exchanged the chloride counterions for hexafluorophosphate in water, which indeed provided us with homogeneous solutions in DMF, DMSO, and DMAc. GPC in DMF at 80 °C proved disappointing since we were not able to establish satisfactory chromatographic conditions to suppress undesirable interactions of the polyrotaxane with the column material. At this point further optimization of the eluent is necessary to obtain meaningful data, which once identified will also be used to isolate fractionated polymer samples to facilitate MALDI-TOF analysis.^{133,134} The chemical character of polyrotaxane **19** is certainly different from that of cucurbit[6]uril before and after counterion exchange. All polyrotaxanes isolated before counterion exchange were soluble in water, in which **1** is not.⁸⁹ A solvent other than water “doped” with high levels of protons or alkali-metal ion salts has so far not been identified for cucurbit[6]uril.

With our success in triggering cucurbit[6]uril to catalyze the cycloaddition between monomers which act as stopper groups, we had an opportunity to unravel the polymerization mechanism further by studying the catalytic behavior of **1** in the presence of a matched pair of monomers, only one of which containing a stopper group (Figure 17). Monomers **6** and **14** were reacted under a set of conditions summarized in Table 4.

The maximum reaction progress achieved was the formation of a single triazole ring for every pair of monomers. This shows how vital it was to have selected monomers which cannot be encapsulated by cucurbit[6]uril. Figure 17 is an attempt to translate our observations into a mechanistic scheme. The formation of [3]semirotaxane **24** (Figure 17) in the presence of 2 equiv of **1** (Table 4, entries A and B) suggests the encapsulation of monomer **6** through **1** is the first step, being faster than the catalytic threading pathway to give **24**. The fact that the unstoppered monomers **5** and **6** were completely unreactive under almost identical conditions supports this view (Table 1, entry B). A third equivalent of cucurbit[6]uril is therefore required to complex to monomer **14** and subsequently to form the catalytically active ternary complex, which is only possible because the first equivalent of cucurbit[6]uril encapsulating monomer **6** can slip to the right or unthread completely. Using a 50% excess of cucurbit[6]uril (Table 4, entry C), catalysis has been suppressed

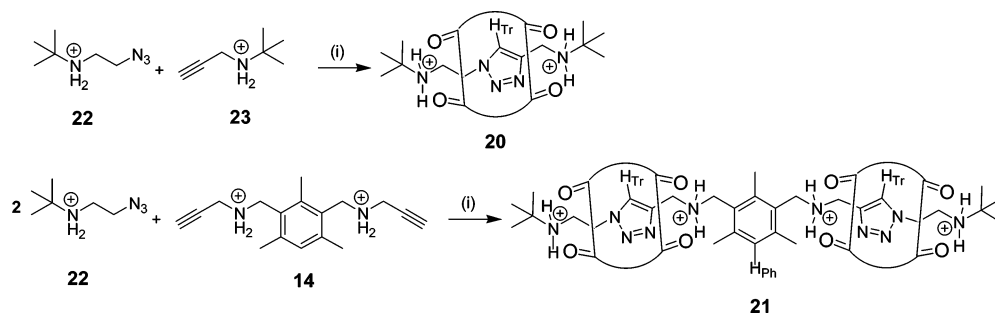


Figure 15. Synthesis of model [2]- and [3]rotaxanes **20** and **21**. Conditions and reagents: (i) **1**, 6 N HCl, 72 h, room temperature, 71% (**20**) and 88% (**21**).

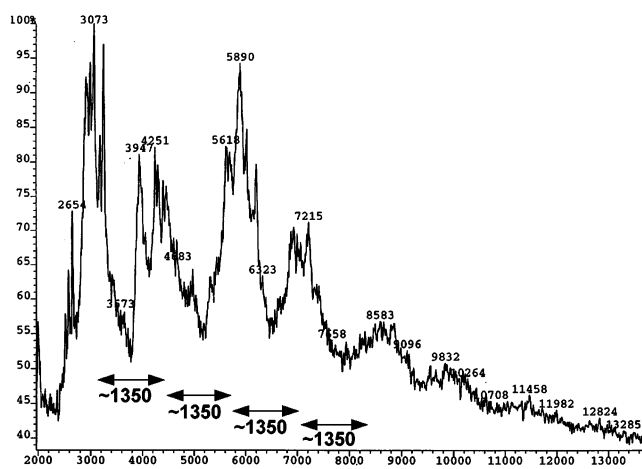


Figure 16. A representative MALDI-TOF mass spectrum for polyrotaxane **19** (Table 3, entry A).

by shifting the equilibrium of threaded to unthreaded monomer **6** strongly to the side of the former, thus significantly reducing the statistical change for the formation of the catalytic ensemble. This reading of the data is in line with entries D and E (Table 4), indicating that longer reaction times or elevated temperatures counteract the shift in equilibrium caused by the presence of an excess of **1**.

6. Branched Polyrotaxanes. Encouraged by our ability to produce catalytically self-threading main-chain polyrotaxanes, we investigated the possibility of extending our architectural repertoire to branched analogues. For simplicity we opted for an A_2/B_3 system.^{135,136} The A_2 component was already in hand in the guise of **18**. Since 2,4,6-tris(bromomethyl)mesitylene is commercially available, we prepared the B_3 monomer **25** responsible for the branching analogously to the synthesis of **14**. *N*-{2,4,6-Trimethyl-3,5-bis[(2-propynyl-amino)methyl]benzyl}-2-propyn-1-amine trihydrochloride (**25**) was obtained in 88% yield. Polymerizations were carried out using different ratios of **1**, **18**, and **25** as shown in Table 5, with the structure of the expected branched polyrotaxane **26** illustrated in Figure 18.

The degree of polymerization for branched polyrotaxane **26** was determined by comparing the integral of the triazole proton to that of the phenyl proton in the same way as already described in detail for the linear main-chain polyrotaxanes. The same problems were encountered as for the linear analogues when attempts were made to obtain molecular weight information through MALDI-TOF mass spectrometry and GPC. Figure 19 shows a representative MALDI-TOF spectrum (for Table 5, entry B) with features similar to those discussed earlier. The most striking patterns are once

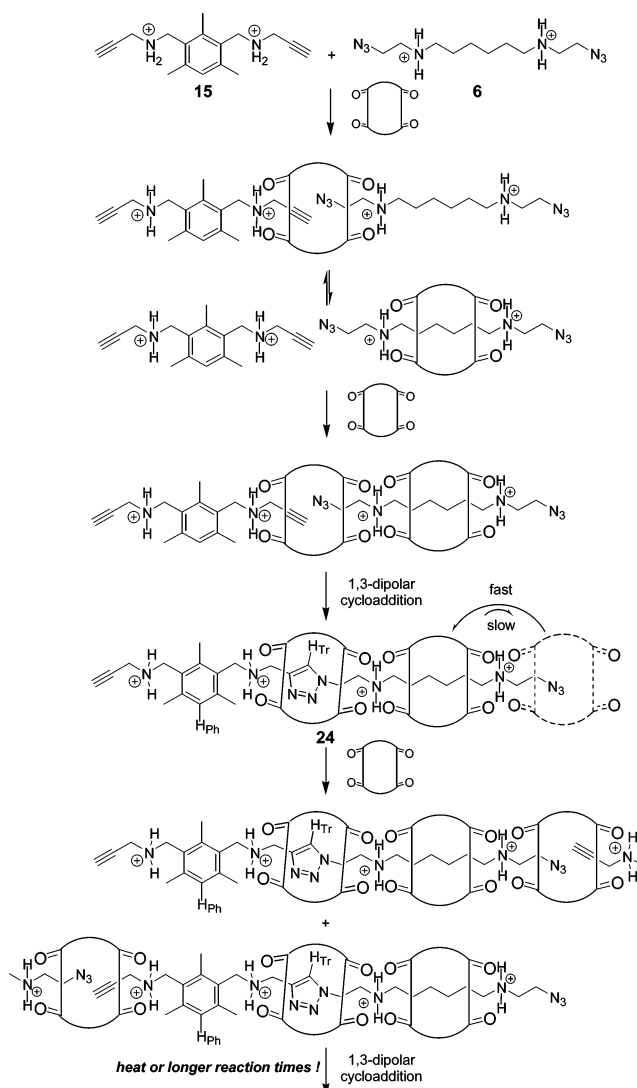
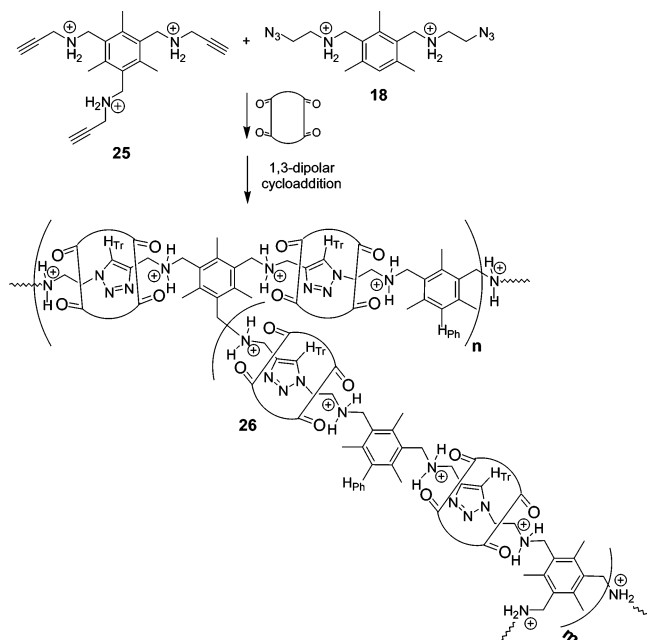
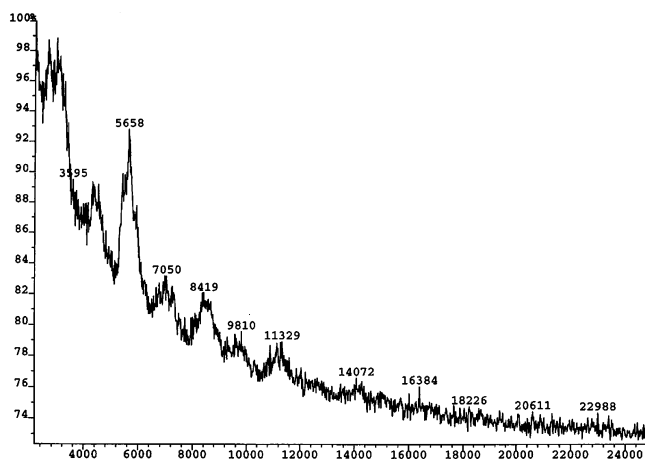


Figure 17. Formation of [3]semirotaxane **24** and attempted polymerization of monomers **6** and **14** in the presence of **1**. See Table 4 for reaction conditions.

again a wide mass range of molecular ions and the separation of the clusters corresponding approximately to the molar mass of half a threaded repeat unit. ^1H and ^{13}C NMR analysis is the obvious tool to determine the degree of branching of polyrotaxane **26**. In our attempts to quantitatively analyze these spectra, we encountered broad, and most of the time featureless, resonances, an unsurprising consequence of the polymeric nature of the molecules as well as the restricted motions of the macrocyclic cavitand. This made it essentially impossible to integrate and assign reso-

Table 5. Reaction Conditions and Molecular Weight Data for the Polymerization of Monomers **18 and **25** in the Presence of **1** Dissolved in 6 N HCl**

entry	[18]/molar equiv	[25]/molar equiv	[1]/molar equiv	t/h	T/°C	yield/%	M_n (^1H NMR)	DP_n^a
A	1.00	1.00	2.00	96 + 24	20 + 60	80	5000	2
B	1.50	1.00	3.00	72	20	70	22000	8
C	1.50	1.00	3.00	96	20	75	32000	13
D	1.00	1.00	3.00	144	60	78	34000	14

^a DP_n = degree of polymerization.**Figure 18.** Synthesis of branched/hyperbranched polyrotaxane **26** via catalytic self-threading. See Table 5 for reaction conditions. Conditions and reagents for the synthesis of **25**: (i) propargylamine (neat), 0 °C $\rightarrow$ room temperature, 60 °C, 16 h, 100% (crude); (ii) 1 N HCl, in Et₂O, room temperature, 88%.**Figure 19.** A representative MALDI-TOF mass spectrum for polyrotaxane **26** (Table 5, entry B).

nances reliably. Furthermore, the probably most promising approach to calculate the degree of branching, via end group analysis of unreacted alkyne and azide groups in conjunction with the corresponding benzyl resonances of monomer **25**, suffered from overlapping signals such as the benzyl resonances of **18**.

Conclusions

Catalytic self-threading has been established as a new route for the synthesis of polyrotaxanes. Investigations

into the polymerization mechanism have brought to light that the catalyst cucurbit[6]uril is highly sensitive toward monomer structure, with a poorly designed monomer resulting in complete inactivity. By using monomers which contain stopper groups to prevent cucurbit[6]uril from becoming deactivated through threading, we were able to prepare main chain polyrotaxanes in high yields with molecular weights up to 39000. These polyrotaxanes are structurally perfect in the sense that exactly two cucurbit[6]uril molecules are threaded onto each structural repeat unit, though still at the cost of restricting severely the movements of the macrocycle along the polymer backbone. The concept was extended to the synthesis of branched polyrotaxanes, though more detailed structural analysis is necessary to establish the degree of branching for these polymers.

Experimental Section

All manipulations of air- and moisture-sensitive compounds were performed under an atmosphere of nitrogen. NMR solvents were obtained commercially and used as received unless stated otherwise: *d*₄-methanol (Aldrich Chemicals), *d*₆-DMSO (Aldrich Chemicals), *d*₂-water (Apollo Scientific), *d*₂-sulfuric acid (Apollo Scientific), DCl (35% aqueous solution (w/w)) (Aldrich Chemicals), and DCl (20% aqueous solution (w/w)) (Apollo Scientific).

Solvents were dried by prolonged reflux over a suitable drying agent (in parentheses). They were distilled and degassed prior to use: tetrahydrofuran (sodium metal and benzophenone), dichloromethane (calcium hydride), diethyl ether (lithium aluminum hydride), ethanol (4 Å molecular sieves), chloroform (P₂O₅). Reagents were obtained commercially at the highest purity available and were used as received unless stated otherwise: 1,6-diaminohexane, 1,6-dichlorohexane, ethanolamine, nylon 6/6, glycoluril, propargylamine, propargyl bromide (in toluene), *tert*-butyl amine, 2-(*N*-*tert*-butylamino)ethanol, and 2,4,6-tris(bromomethyl)-mesitylene (Aldrich Chemicals), thionyl chloride (Fluka), sodium azide (Avocado), 2,4-bis(chloromethyl)-1,3,5-trimethylbenzene (Lancaster).

*N*¹,*N*⁶-Bis(2-hydroxyethyl)-1,6-hexanediamine¹³⁷ (**2**), *N*¹,*N*⁶-bis(2-propynyl)-1,6-hexanediamine dihydrochloride¹²⁷ (**6**), *N*¹,*N*⁶-bis(*tert*-butoxycarbonyl)-1,6-hexanediamine¹²⁶ (**7**), poly(iminohexamethylene)¹²⁸ (**11**) (GPC (CHCl₃): M_n = 20000, M_w = 55000, DP_n = 202)), *N*-(*tert*-butyl)azidoethylamine hydrochloride^{86,138,139} (**22**), and *N*-(*tert*-butyl)propargylamine hydrochloride¹⁴⁰ (**23**) were prepared according to published syntheses.

NMR solvents were obtained commercially and used as received unless stated otherwise. Infrared spectra were recorded on a Perkin-Elmer 1710 and 1725 series FTIR spectrometer using KBr plates. Absorptions are abbreviated as follows: vs (very strong), s (strong), m (medium), w (weak), br (broad), sh (shoulder). NMR spectra were recorded on Bruker DRX400 (400 MHz, ^1H ; 100 MHz, ^{13}C) and Bruker AC250 (250 MHz, ^1H ; 62.5 MHz, ^{13}C) instruments.

MALDI-TOF mass spectrometry analysis was carried out by the ULIRS (University of London Intercollegiate Research Service) on a Fisons VG ToFSpec instrument using a semiautomatic probe, holding 12 samples at a time. The instrument was differentially pumped. The pressure in the analyzer region was 10⁻⁸ Torr, and the source region pressure was 10⁻⁷ Torr.

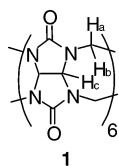
The samples were prepared in water using a 1000-fold molar excess of the matrix (indoleacrylic acid). Typically 1 μ L of sample matrix mixture was deposited onto the sample position of a stainless steel target. The MALDI used a nitrogen laser at 337 nm (UV) with a shot frequency of 10 Hz and was operated at 28 kV accelerating voltage. Typically, 100 shots were acquired and averaged to produce the spectrum.

GPC analysis in DMF was performed by RAPRA on a Polymer Laboratories GPC-210 using DMF with ammonium acetate as eluent. Columns used were two 10 μ m Plgel mixed bed B columns (30 cm; 5×10^2 to 10×10^6 Da), with a flow rate of 1.0 mL/min (nominal) at 80 °C. The system was calibrated with PMMA standards. GPC analysis in CHCl_3 (stabilized with ethanol) as the eluent was performed by RAPRA on a Viscotek system equipped with a refractive index detector (also Viscotek). Columns used were two 10 μ m Plgel mixed bed B columns (30 cm; 5×10^2 to 10×10^6 Da) with a flow rate of 1.0 mL/min (nominal) at 40 °C. The system was calibrated with PS standards. Elemental analyses were carried out at the Analytical Services of the University of North London.

Cucurbituril 1. Glycoluril (7.0 g, 49 mmol), formaldehyde (10.5 mL of a 37% w/w aqueous solution, 150 mmol), concentrated HCl (16 mL), and water (35 mL) were placed in a round-bottom flask and heated to reflux until all solid had dissolved. A few minutes later, the reaction mixture turned cloudy and a precipitate was formed. The contents of the flask were poured into 350 mL of ice-cooled water, and a white powder precipitated. This was filtered off and subsequently washed with water (10 mL), ethanol (10 mL), and diethyl ether (10 mL). The white powder was dried in vacuo over P_2O_5 for a week to yield the precursor for **1**. (Longer drying times better the results.) Yield: 5.42 g (42%).

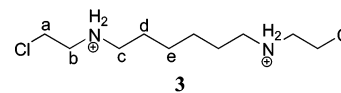
To the well-dried solid was added carefully concentrated H_2SO_4 (2.2 mL of acid/g of solid). The mixture was heated under vigorous stirring to 110–120 °C until all solid had dissolved. Initially, the reaction mixture turned into a brown suspension but subsequently became a brown viscous solution. After being cooled to room temperature, the mixture was poured into ice cold water (22 mL/g of starting material). A small amount of precipitate was filtered rapidly through a sintered glass funnel with suction. The light brown colored filtrate was heated gently to yield beige crystals of **1**.

Crude **1** was dissolved in the minimum amount of formic acid/water (1:1 (v/v)) and heated to reflux for about 1 h. The resulting solution was filtered hot with suction. The filtrate was allowed to cool to room temperature. Water (twice the volume of the filtrate) was added to yield a white powder. The white precipitate was washed with hot water, dried over P_2O_5 under vacuum for one week, and stored over P_2O_5 . Yield: 1.81 g (21%). Mp: >300 °C. IR (KBr, Nujol mull, cm^{-1}): 3469 (NH–), 2998 (C–H), 1738 (C=O). ^1H NMR (250 MHz, DCl): δ 4.46 (d, 12H, $^2J_{\text{HH}} = 15.6$ Hz, H_b), 5.54 (d, 12H, $^2J_{\text{HH}} = 15.6$ Hz, H_a), 5.75 (s, 12H, H_c). ^{13}C NMR (62.5 MHz, DCl/ D_2O , 35% (w/w)): δ 51.55 (CH_2), 71.60 (CH), 158.1 (C=O). Molar mass of $\text{C}_{36}\text{H}_{36}\text{O}_{12}\text{N}_{24}$: 996.8418. MS (FAB + ve): m/z 997 [$\text{M} + \text{H}$] $^+$, 1019 [$\text{M} + \text{Na}$] $^+$, 1035 [$\text{M} + \text{K}$] $^+$. Anal. Calcd for $\text{C}_{36}\text{H}_{36}\text{O}_{12}\text{N}_{24} \cdot 2.5\text{H}_2\text{O}$: C, 41.50; H, 3.96; N, 32.26. Found: C, 41.38; H, 3.93; N, 31.86.



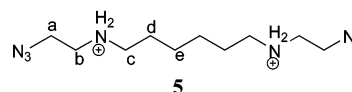
N^1, N^6 -Bis(2-azidoethyl)-1,6-hexanediamine Dihydrochloride (5). **2** (2.50 g, 12.2 mmol) and thionyl chloride (25 mL) were heated to reflux for 100 min. The reaction mixture was allowed to cool to room temperature before excess thionyl chloride was removed under reduced pressure. The resulting dark brown residue was washed with 2-propanol. The dark colored powder was filtered off and recrystallized from 2-pro-

panol to yield **3** as a light brown solid, used without further purification. Yield (crude): 2.5 g (65%). ^1H NMR (250 MHz, D_2O): δ 1.40, (m, 4H, e), 1.65 (m, 4H, d), 3.05 (t, 4H, $^3J_{\text{HH}} = 7.1$ Hz, c), 3.45 (t, 4H, $^3J_{\text{HH}} = 5.4$ Hz, a), 3.80 (t, 4H, $^3J_{\text{HH}} = 5.4$ Hz, b).



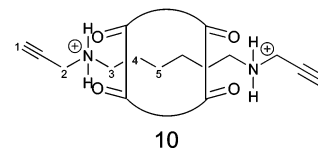
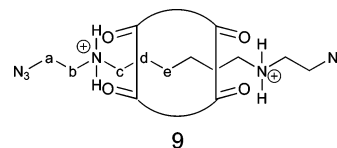
Compound **3** (1.73 g, 5.54 mmol) and sodium azide (1.00 g, 15.4 mmol) were dissolved in water (25 mL). The light yellow solution was heated to 75 °C under vigorous stirring overnight. Sodium hydroxide pellets (0.50 g, 12.5 mmol) were added, and the solution was stirred for 1 h. The precipitate was filtered off and the filtrate extracted with diethyl ether (4×25 mL). The combined organic layers were washed with saturated NaCl solution (30 mL) and dried over anhydrous MgSO_4 . N^1, N^6 -Bis(2-azidoethyl)-1,6-hexanediamine (**4**) was isolated as a pale yellow oil after removal of the solvent under reduced pressure. Yield (crude): 0.66 g (60%).

The oil was dissolved in ethanol (5 mL), and a 1 N solution of HCl in ether was added dropwise. After 3 h of stirring at room temperature the precipitated white solid was filtered off. Recrystallization from ethanol/toluene (50:50 (v/v)) yielded light yellow, microcrystalline **5**. Yield: 0.75 (84%). IR (Nujol mull, KBr, cm^{-1}): 3585 (w), 3392 (m), 3158 (m), 2438 (m), 2105 (s), 2088 (vs). ^1H NMR (250 MHz, D_2O): δ 1.35 (m, 4H, e), 1.65 (m, 4H, d), 3.05 (t, 4H, $^3J_{\text{HH}} = 7.1$ Hz, c), 3.15 (t, 4H, 5.4 Hz, a), 3.75 (t, 4H, $^3J_{\text{HH}} = 5.4$ Hz, b). ^{13}C NMR (100 MHz, D_2O): δ 28.07 (e), 28.11 (d), 49.08 (c), 49.76 (a), 50.39 (b). Molar mass of $\text{C}_{10}\text{H}_{24}\text{Cl}_2\text{N}_8$: 327.26. MS (EI): m/z 256 ($\text{M} - 2\text{Cl}$). Anal. Calcd: C, 36.70; H, 7.39; N, 34.24. Found: C, 36.81; H, 7.52; N, 34.50.



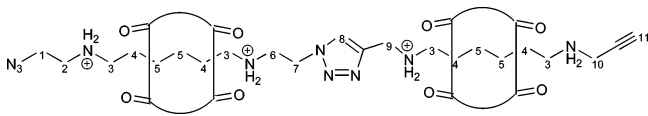
Attempted Polymerization of Monomers 5 and 6 (See Table 1). [Entry A (Table 1)] **1** (32.2 mg, 31.0 μ mol) was dissolved in DCl (1 mL, 20 wt % DCl in D_2O) and the solution stirred for 30 min. **5** (4.9 mg, 16 μ mol) was added, and the solution was stirred vigorously at room temperature for another 10 min before the addition of alkyne **6** (4.0 mg, 16 μ mol). The resulting solution was stirred at 20 °C for 24 h. The reaction was monitored by ^1H NMR.

Spectroscopic Data. [Entries A–C (Table 1), pseudotaxanes **9** and **10**] ^1H NMR (250 MHz, D_2O): δ 0.45 (m, 8H, 5 + e), 0.75 (m, 8H, 4 + d), 2.95 (t, 8H, $^3J_{\text{HH}} = 6.5$ Hz, 3 + c), 3.15 (t, 2H, $^2J_{\text{HH}} = 2.5$ Hz, 1), 3.35 (t, 4H, $^3J_{\text{HH}} = 5.5$ Hz, a), 3.85 (t, 4H, $^3J_{\text{HH}} = 5.5$ Hz, b), 4.15 (d, 2H, $^2J_{\text{HH}} = 2.5$ Hz, 2), 4.43 (d, 24H, QQ), 5.55 (s, 12H, QQ), 5.75 (d, 12H, QQ). ^{13}C NMR (62.5 MHz, D_2O): δ = 27.99 (5), 28.06 (e), 28.67 (4), 29.16 (d), 39.13 (a), 49.07 (c), 49.43 (3), 50.38 (b), 50.59 (2), 54.38 (QQ), 70.98 (–CCH), 73.26 (QQ), 76.02 (1), 159.19 (QQ).

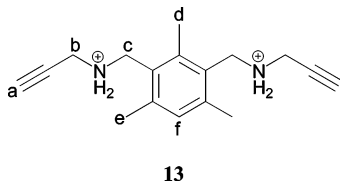


The stoichiometry between triazole protons and monomer repeat units was determined by comparing the integral of H_8 with those of $\text{H}_4 + \text{H}_5$, [Entry D (Table 1)] ^1H NMR (250 MHz,

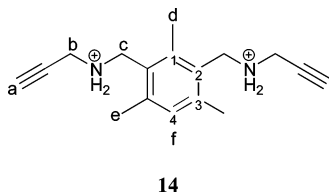
D₂O): δ 0.44 (br s, 8H, 5), 0.70 (br s, 8H, 4), 8.55 (s, 1H, 8). [Entry E (Table 1)] ¹H NMR (250 MHz, D₂O): δ 0.46 (br m, 8H, 5), 0.72 (br m, 8H, 4), 8.51 (s, 0.7H, 8). [Entry F (Table 1)] ¹H NMR (250 MHz, D₂O): δ 0.48 (br s, 8H, 5), 0.75 (br s, 8H, 4), 8.51 (s, 0.8 H, 8).



N-{2,4,6-Trimethyl-3-[(2-propynylamino)methyl]benzyl}-2-propyn-1-amine Dihydrochloride (14). **12** (0.51 g, 2.3 mmol) was added to an excess of propargylamine (2.0 mL, 40.6 mmol) at 0 °C. When addition was complete, the light brown solution started to solidify within 5 min. The solid was heated to redissolve, then refluxed for 4 h, and finally left to stir overnight at room temperature. Excess propargylamine was distilled off at ambient pressure. The remaining brown residue was dissolved in water (5 mL), sodium hydroxide pellets (0.2 g, 5.0 mmol) were added, and the suspension was stirred for 2–3 h. The aqueous layer was extracted with chloroform (4 × 5 mL). The combined organic layers were washed with saturated NaCl solution (10 mL). The organic phase was dried over magnesium sulfate and filtered, and the solvent was removed under reduced pressure, yielding beige solid **13**. Yield (crude): 0.55 g (100%). ¹H NMR (250 MHz, CDCl₃): δ 2.29 (t, 2H, ⁴J_{HH} = 2.5 Hz, a), 2.35 (s, 6H, e), 2.45 (s, 3H, d), 3.49 (d, 4H, ⁴J_{HH} = 2.5 Hz, b), 3.84 (s, 4H, c), 6.86 (s, 1H, f).

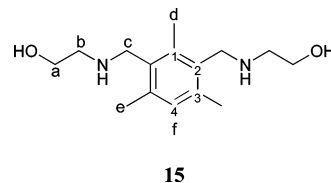


Crude **13** (0.55 g) was dissolved in chloroform (5 mL), to which a 1 N solution of HCl in diethyl ether (4.5 mL) was added dropwise. A dark yellow precipitate was formed immediately, and the suspension was stirred overnight. The precipitate was filtered and washed with chloroform (5 mL) followed by diethyl ether (10 mL). Recrystallization from ethanol/toluene (80:20 (v/v)) gave **14** as a light yellow powder. Yield: 0.60 g (86%). ¹H NMR (250 MHz, D₂O): δ 2.30 (s, 6H, e), 2.40 (s, 3H, d), 3.05 (t, 2H, ⁴J_{HH} = 2.5 Hz, a), 4.09 (t, 4H, ⁴J_{HH} = 2.5 Hz, b), 4.44 (s, 4H, c), 7.15 (s, 1H, f). ¹³C NMR (100 MHz, D₂O, DSS): δ 18.43 (d), 22.13 (e), 39.80 (b), 47.87 (c), 76.06 (a), 81.46 (–CCH), 129.52 (4), 134.39 (1), 141.45 (2), 143.62 (3). Molar mass of C₁₇H₂₄Cl₂N₂: 327.29. MS (FAB + ve): m/z 255 [M – 2HCl]⁺, 291 [M – HCl]⁺. Anal. Calcd for C₁₇H₂₆N₂Cl₂O: C, 59.13; H, 7.59; N, 8.62. Found: C, 59.46; H, 7.61; N, 8.79.

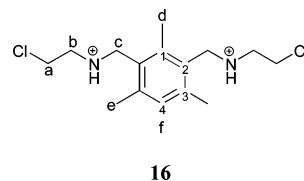


N-(2-Hydroxyethyl)-N-(3-[(2-hydroxyethyl)aminol]-methyl)-2,4,6-trimethylbenzylamine (15). **12** (1.0 g, 4.6 mmol) was added portionwise to ethanolamine (5.6 mL, 92 mmol) over a period of 20 min under vigorous stirring at 120–130 °C. The mixture was heated at 150–160 °C for 6 h, allowed to cool to room temperature, and then treated with a 0.05 N methanolic solution of NaOH (20 mL). Then solution was cooled to 0 °C and stirred for several hours at this temperature before the precipitated sodium chloride was filtered off. First methanol was removed from the filtrate under reduced pres-

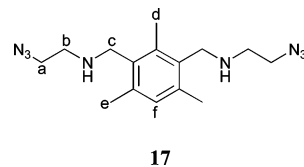
sure (20 mmHg) at room temperature, and then excess ethanolamine was distilled off at 50 °C (0.8 mmHg). The remaining white solid was recrystallized from ethanol to yield **15** as a white powder. Yield: 0.80 g (66%). ¹H NMR (250 MHz, D₂O): δ 2.25 (s, 6H, e), 2.30 (s, 3H, d), 2.80 (t, 4H, ³J_{HH} = 7.5 Hz, b), 3.54 (t, 4H, ³J_{HH} = 7.5 Hz, a), 3.77 (s, 4H, c), 6.95 (s, 1H, f). ¹³C NMR (100 MHz, D₂O): δ 18.64 (e), 22.24 (d), 41.77 (c), 48.23 (b), 51.83 (a), 129.43 (4), 134.44 (1), 141.50 (3), 143.60 (2). Molar mass of C₁₅H₂₆N₂O₂: 266.38. MS (FAB + ve): m/z 267 [M + H]⁺.



N-(2-Chloroethyl)-N-(3-[(2-chloroethyl)amino]methyl)-2,4,6-trimethylbenzylamine Dihydrochloride (16). **15** (2.22 g, 8.37 mmol) was dissolved in chloroform (25 mL) and the solution cooled to –10 °C. A solution of thionyl chloride (2.50 mL, 34.0 mmol) in chloroform (10 mL) was added dropwise under vigorous stirring over a period of 15 min. The resulting white suspension was allowed to warm to room temperature and then was heated to reflux for 3 h. During this time a light yellow precipitate formed. The mixture was allowed to cool to room temperature before methanol (10 mL) was added, and the reaction mixture was stirred for a further 15 min. Removal of the solvent under reduced pressure yielded a light yellow solid. Recrystallization from ethanol/acetone (60:40 (v/v)) gave **16** as a white powder. Yield: 1.76 g (70%). ¹H NMR (250 MHz, D₂O): δ 2.37 (s, 6H, e + f), 2.43 (s, 3H, d), 3.58 (t, 4H, ³J_{HH} = 5.5 Hz, b), 3.91 (t, 4H, ³J_{HH} = 5.5 Hz, a), 4.55 (s, 4H, c), 7.14 (s, 1H, g). ¹³C NMR (100 MHz, D₂O): δ 18.36 (e), 22.10 (d), 48.18 (a), 52.08 (c), 59.08 (b), 129.88 (4), 134.28 (1), 141.28 (3), 143.29 (2). Molar mass of C₁₅H₂₆N₂Cl₄: 374.09. MS (EI): m/z 302 [M – 2HCl].

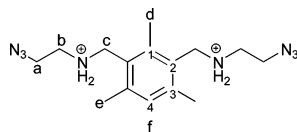


N-(2-Azidoethyl)-N-(3-[(2-azidoethyl)amino]methyl)-2,4,6-trimethylbenzylamine Dihydrochloride (18). Compound **16** (1.39 g, 4.60 mmol) and sodium azide (1.19 g, 18.4 mmol) were dissolved in water (50 mL). The solution was heated at 75 °C overnight. Sodium hydroxide pellets (0.81 g, 20 mmol) were added neat, and the solution was stirred for an additional 2 h. A light yellow oil formed which was extracted into chloroform (4 × 30 mL). The organic phase was washed with a saturated aqueous solution of NaCl (30 mL) and dried over anhydrous MgSO₄. Diazide **17** was obtained as a pale yellow oil after the solvent was removed under reduced pressure. Yield (crude): 1.02 (87%). IR (film, KBr, cm^{–1}): 3328 (br), 3006 (s), 2922 (s), 2101 (vs), 1667 (m), 1451 (s). ¹H NMR (250 MHz, CDCl₃): δ 2.33 (s, 6H, e), 2.42 (s, 3H, d), 2.90 (t, 4H, ³J_{HH} = 5.6 Hz, b), 3.46 (t, 4H, ³J_{HH} = 5.6 Hz, a), 3.77 (s, 4H, c), 6.95 (s, 1H, f).

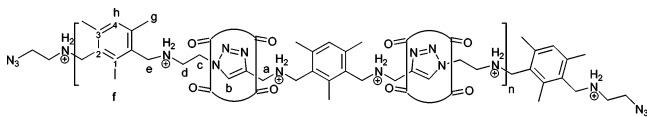


Compound **17** was dissolved in anhydrous methanol (5 mL). A solution of 1 N anhydrous HCl in diethyl ether (6.5 mL) was

added dropwise under vigorous stirring at 0 °C over a period of 20 min. The reaction mixture was stirred for a further 5 h, before the solvent was removed under reduced pressure. The remaining solid was recrystallized from ethanol to yield **18** as a pale yellow microcrystalline powder. Yield: 1.20 g (96%). ¹H NMR (250 MHz, D₂O): δ 2.32 (s, 6H, e), 2.42 (s, 3H, d), 3.36 (t, 4H, ³J_{HH} = 5.6 Hz, b), 3.81 (t, 4H, ³J_{HH} = 5.6 Hz, a), 4.39 (s, 4H, c), 7.15 (s, 1H, f). ¹³C NMR (100 MHz, D₂O): δ 18.50 (d), 22.15 (e), 48.32 (c), 49.28 (b), 49.38 (a), 129.54 (4), 134.37 (1), 141.36 (3), 143.46 (2). Molar mass of C₁₅H₂₆N₈Cl₂: 389.33. MS (FAB + ve): *m/z* 317 [M - 2HCl]⁺, 353 [M - HCl]⁺. Anal. Calcd for C₁₅H₂₆N₈Cl₂: C, 46.28; H, 6.73; N, 28.78. Found: C, 46.26; H, 6.71; N, 28.70.

**18**

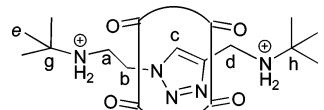
General Procedure for the Synthesis of Polyrotaxane 19. (Example, Table 3, entry A) **1** (200 mg, 200 μmol) was dissolved in 6 N HCl (4 mL) and the solution stirred for 30 min. Azide **18** (39.1 mg, 100 μmol) was added under vigorous stirring at room temperature. After 10 min alkyne **14** (32.3 mg, 100 μmol) was added, upon which the suspension became homogeneous. The resulting solution was stirred at room temperature for 48 h (see Table 3 for reaction conditions, yield, and molar mass data (for the latter see the text for the calculation procedure)). The solution was precipitated into a large excess of acetone/ethanol (50:50, v/v) to yield a white solid, which became pale yellow during filtration. To remove excess **1**, the solid was dissolved in hot water (~80 °C) and the solution stirred for 2 h. Undissolved material was filtered off. The solvent was removed under reduced pressure to a yield colorless film, which was dried in vacuo over P₂O₅ for one week. Yield: 185 mg (68%). (Spectroscopic data for Table 3, entries A–P) ¹H NMR (400 MHz, D₂O): δ 2.75 (s, 6H, g), 3.35 (s, 3H, f), 3.81 (br, 2H, d), 4.02 (br, 2H, c), 4.31 (t, 12H, QQ), 4.52 (s, 2H, e), 4.65 (s, 2H, a), 5.55–5.95 (m, 24H, QQ), 6.54 (s, XH, b), 7.21 (s, 1H, h). ¹³C NMR (100 MHz, D₂O + DSS): δ 18.96 (g), 19.09 (g), 22.65 (f), 22.77 (f), 42.63–42.84 (d), 43.17–43.39 (c), 49.22–49.45 (a), 51.05–51.28 (e), 54.15 (QQ), 54.33 (QQ), 73.16 (QQ), 122.73 (Tr, =CH), 129.86 (4), 130.17 (2), 133.01 (1), 142.17 (3), 143.52 (Tr, =CR), 159.28–159.38 (QQ).

**19**

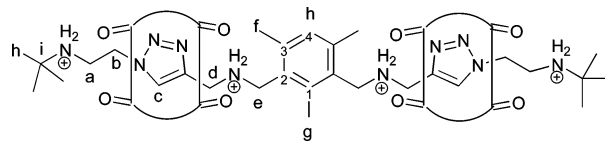
Counterion Exchange of Polyrotaxane 19. Polyrotaxane **19** (Table 3, entry F) (50 mg) was dissolved in water (3 mL), to which a saturated solution of ammonium hexafluorophosphate was added dropwise. Upon addition the clear solution turned cloudy. Addition continued until no further precipitate was formed. Then the suspension was stirred at room temperature for 3–4 h. Filtration proved too difficult; therefore, the solvent was removed under reduced pressure, and the resulting solid was triturated with water (3 mL) to remove excess salt. The resulting pale yellow sticky solid was dried in vacuo. Counterion-exchanged polyrotaxane **19** was soluble in DMF, DMSO, and DMF.

[2]Rotaxane 20. 1 (200.0 mg, 200.0 μmol) was dissolved in 6 N HCl (5 mL) and the solution stirred at room temperature for 30 min. **22** (36.0 mg, 200 μmol) was added under vigorous stirring followed by **23** (30.0 mg, 200 μmol). The clear solution was stirred at room temperature for 72 h. After removal of the solvent under reduced pressure, the remaining white solid was suspended in hot water (5 mL, ~80 °C) for 1 h. Undissolved cucurbituril was removed by filtration. The

filtrate was precipitated into a large excess of acetone, upon which a white precipitate was formed. It was filtered off, washed with acetone (10 mL), and dried in vacuo over P₂O₅ for one week to afford **20**. Yield (**20** + 9H₂O): 190 mg (71%). ¹H NMR (250 MHz, D₂O): δ 1.61 (s, 9H, e or f), 1.64 (s, 9H, e or f), 3.56 (t, 2H, ³J_{HH} = 6.3 Hz, a), 3.87 (t, 2H, ³J_{HH} = 6.3 Hz, b), 4.26 (dd, 12H, ²J_{HH} = 15.6 Hz, QQ), 4.26 (s, 2H, d), 5.49 (s, 12H, QQ), 5.71 (dd, 12H, ²J_{HH} = 15.6 Hz, QQ), 6.50 (s, 1H, c). ¹³C NMR (100 MHz, D₂O, DSS): δ 28.40 (e or f), 28.86 (e or f), 39.47 (a), 43.41 (b), 50.33 (d), 54.34 (QQ), 61.23 (C(CH₃)₃, g or h), 62.12 (C(CH₃)₃, g or h), 73.20 (QQ), 124.29 (Tr, =CH), 142.65 (Tr, =CR), 159.54 (QQ). Molar mass of C₁₃H₂₉N₅Cl₂·(C₃₆H₃₆O₁₂N₂₄): 1323.16. FAB + ve MS *m/z*: 1251 [M - 2Cl]⁺, 1287 [M - Cl]⁺. Elemental analysis of C₁₃H₂₉N₅Cl₂·(C₃₆H₃₆O₁₂N₂₄)·9H₂O. Calcd: C, 40.24; H, 5.16; N, 27.81. Found: C, 40.11; H, 5.56; N, 27.68.

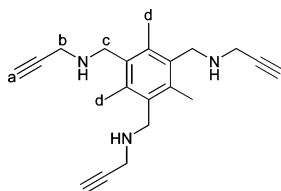
**20**

[3]Rotaxane 21. 1 (150.0 mg, 151.0 μmol) was dissolved in 6 N HCl (4 mL), and **22** (27.0 mg, 151 μmol) was added under vigorous stirring, followed by **14** (21.9 mg, 75.0 μmol). The clear solution was stirred at room temperature for 72 h. The solvent was removed under reduced pressure to yield a light yellow film, which was scraped off. The resulting powder was suspended in methanol and the suspension stirred at room temperature for 3 h. The undissolved residue was filtered off, washed with methanol (3 mL) and acetone (5 mL), and dried in vacuo. The isolated solid was suspended in hot water (5 mL) at 80 °C and the suspension stirred for 1 h before it was filtered to remove excess cucurbituril. The solvent was removed under reduced pressure to yield **21** as an off-white powder, which was dried in vacuo over P₂O₅ for a week. Yield (**21** + 21H₂O): 175 mg (88%). ¹H NMR (250 MHz, D₂O): δ 1.68 (s, 18H, h), 2.68 (s, 6H, f), 2.98 (s, 3H, g), 3.81 (t, 4H, ³J_{HH} = 8.0 Hz, a), 4.03 (t, 4H, ³J_{HH} = 8.0 Hz, b), 4.30 (dd, 24 H, QQ + s, 4H, d), 4.85 (s, 4H, e), 5.59 (s, 24H, QQ), 5.73 (t, 24H, QQ), 6.54 (s, 2H, c), 7.29 (s, 1H, h). ¹³C NMR (100 MHz, D₂O): δ 20.54 (g), 29.16 (h), 42.91 (a), 42.83 (b), 49.17 (d), 50.67 (e), 54.32 (QQ), 54.50 (QQ), 62.74 (i), 72.91 (QQ), 73.03 (QQ), 124.05 (Tr, =CH), 130.22 (Ar, C₄), 134.49 (Ar, C₁), 141.86 (Ar, C₂ + C₆), 142.75 (Ar, C₃ + C₅), 143.28 (Tr, =CR), 159.36 (QQ), 159.71 (QQ). Molar mass of C₂₉H₅₄N₁₀Cl₄·(C₃₆H₃₆O₁₂N₂₄)₂: 2678.32. MS (ES): *m/z* 2536.50 [M - 4Cl]⁺, 2573.10 [M - 3Cl]⁺. Anal. Calcd for C₂₉H₅₄N₁₀Cl₄·(C₃₆H₃₆O₁₂N₂₄)₂·21H₂O: C, 39.42; H, 5.39; N, 26.91. Found: C, 39.69; H, 5.54; N, 26.58.

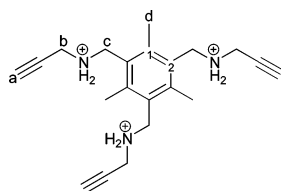
**21**

N-{2,4,6-Trimethyl-3,5-bis[(2-propynylamino)methyl]benzyl}-2-propyn-1-amine Trihydrochloride (25). 2,4,6-Tris(bromomethyl)mesitylene (1.0 g, 2.5 mmol) was added portionwise to propargylamine (3.4 mL, 50 mmol) at 0 °C over a period of 30 min. A dark red solid was obtained, which was allowed to warm to room temperature and subsequently heated to reflux for 24 h. After excess propargylamine was distilled off at ambient pressure, the dark red residue was suspended in chloroform (5 mL), and 1 N aqueous NaOH (8 mL) was added. The layers were separated after the suspension was stirred for 4 h. The aqueous phase was extracted with CHCl₃ (4 × 5 mL). The combined organic layers were washed with saturated aqueous NaCl solution and dried over anhydrous MgSO₄. The solvent was removed from the filtrate under

reduced pressure to yield the triamine as a red brown oily residue. Yield (crude): 0.90 g (100%). ^1H NMR (250 MHz, CDCl_3): δ 1.06 (br, 3H, NH), 2.36 (t, 3H, $^4J_{\text{HH}} = 2.5$ Hz, a), 2.45 (s, 9H, d), 3.50 (d, 6H, $^4J_{\text{HH}} = 2.5$ Hz, b), 3.86 (s, 6H, c).



The crude triamine was dissolved in dry CHCl_3 (5 mL), and a solution of anhydrous 1 N HCl in diethyl ether (11 mL) was added dropwise over a period of 15 min. The brown colored reaction mixture was stirred overnight. A light brown precipitate was filtered off and washed with diethyl ether (10 mL). Recrystallization from ethanol/toluene (80:20 (v/v)) yielded **25** as a light orange solid. Yield: 0.95 g (88%). ^1H NMR (250 MHz, D_2O): δ 2.49 (s, 9H, d), 3.06 (t, 3H, $^4J_{\text{HH}} = 2.5$ Hz, a), 4.04 (t, 6H, $^4J_{\text{HH}} = 2.5$ Hz, b), 4.54 (s, 6H, c). ^{13}C NMR (100 MHz, D_2O): δ 19.53 (d), 39.98 (b), 48.25 (c), 75.99 (a), 81.70 (—CCH), 131.02 (1), 144.06 (2). Molar mass of $\text{C}_{21}\text{H}_{31}\text{N}_3\text{Cl}_4$: 429.15. MS (FAB + ve): m/z 322 $[\text{M} - 3\text{HCl}]^+$, 358 $[\text{M} - 2\text{HCl}]^+$. Anal. Calcd for $\text{C}_{21}\text{H}_{31}\text{N}_3\text{Cl}_4$: C, 58.54; H, 7.02; N, 9.75. Found: C, 58.38; H, 6.93; N, 9.73.



25

General Procedure for the Synthesis of Branched Polyrotaxane 26. (Example, Table 5, entry D) **1** (300 mg, 300 μmol) was dissolved in 6 N HCl (5 mL) under vigorous stirring. Alkyne **25** (43.1 mg, 100 μmol) was added to yield a yellow colored solution, which was stirred at room temperature for 1 h before the addition of azide **18** (38.9 mg, 100 μmol). The resulting solution was heated at 60 $^\circ\text{C}$ for 192 h. Precipitation into acetone/ethanol (25 mL, 50:50 (v/v)) yielded a white solid which became slightly sticky and yellow colored during suction filtration. The solid was dried in vacuo over P_2O_5 for one week. Yield: 308 mg, 78%. See the text for a discussion of ^1H NMR and MALDI-TOF data.

Acknowledgment. We thank the EPSRC for providing an IPSI project studentship for D.T. Thanks are also due to Dr. Welham of the University of London ULIRS service for carrying out MALDI-TOF spectrometry, Dr. Ball for providing us with ES mass spectrometry data, and the EPSRC RAPRA service for GPC characterization data.

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